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under the same temperature a....

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Chapter 11 - Molecular Composition of Gases 11-1 Volume-Mass Relationships of Gases I. Measuring and Comparing the Volumes of Reacting Gases A. Observations of Gay-Lussac 1. 2 liters H<sub>2</sub> + 1 liter O<sub>2</sub> → 2 liters H<sub>2</sub>O vapor 2. 2 volumes H<sub>2</sub> + 1 volume O<sub>2</sub> → 2 volumes H<sub>2</sub>O vapor 3. 1 volume H<sub>2</sub> + 1 volume Cl<sub>2</sub> → 2 volumes HCl 4.

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334 CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol H<sub>2</sub> at STP = 22.4 L Oxygen molecule 1 mol O<sub>2</sub>

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at STP = 22.4 L Carbon dioxide molecule 1 mol CO<sub>2</sub> at STP = 22.4 L

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Chapter 11 - Gases - yazvac - Google Sites The atmosphere contains 78% nitrogen gas, 21% oxygen gas, and 1% other gases by volume. Atmospheric pressure decreases higher up in the atmosphere. This is because there are fewer air particles farther away from Earth's surface. CHAPTER 11 Gases 11 review answers molecular composition of gases.

## **Chapter 11 Review Gases Section 1 Answers**

Answer Key Chapter 11 - Chemistry 2e | OpenStax 1. A solution can vary in composition, while a compound cannot vary in composition. Solutions are homogeneous at the molecular level, while other mixtures are heterogeneous.

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